

Quick Vocabulary

Lesson 1

chemical bond attraction between atoms when electrons are shared, transferred, or pooled

chemical equation description of a reaction using element symbols and chemical formulas

chemical reaction process in which atoms of one or more substances rearrange to form one or more new substances

coefficient number placed in front of an element symbol or chemical formula in an equation

law of conservation of mass states that the total mass of the reactants before a chemical reaction is the same as the total mass of the products after the chemical reaction

product new substance produced by a chemical reaction

reactant starting substance in a chemical reaction

Lesson 2

combustion chemical reaction in which a substance combines with oxygen and releases energy

decomposition chemical reaction in which one compound breaks down and forms two or more substances

double replacement describes a reaction in which the negative ions in two compounds switch places, forming two new compounds

single replacement describes a reaction in which one element replaces another element in a compound

synthesis chemical reaction in which two or more substances combine and form one compound

Quick Vocabulary

Lesson 3

activation energy minimum amount of energy needed to start a chemical reaction

catalyst substance that increases the reaction rate by lowering the activation energy of a reaction

endothermic chemical reactions that absorb thermal energy

enzyme catalyst that speeds up chemical reactions in living cells

exothermic chemical reactions that release thermal energy

inhibitor substance that slows down, or even stops, a chemical reaction