Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period: \_\_\_\_\_\_

8th grade – Chapter 8, Elements and Chemical Bonds

***Quick Vocabulary -* Lesson 1**

**chemical bond** force that holds atoms together

**compound** material that is made up of two or more different kinds of atoms joined by chemical bonds

**electron dot diagram** model that represents valence electrons in an atom as dots around the element’s chemical symbol

**valence electron** outermost electron of an atom and participates in chemical bonding

**Lesson 2**

**bond** force that holds atoms together in a compound

**chemical formula** group of chemical symbols and numbers that represent elements and numbers of atoms of each element that make up a compound

**covalent bond** forms when two atoms share one or more pairs of valence electrons

**molecule** group of atoms held together by covalent bonding

**polar molecule** has a slight positive end and a slight negative end because of unequal sharing of electrons

**Lesson 3**

**conduct** to serve as a medium through which something can flow

**ion** atom that is no longer electrically neutral because it has lost or gained valence electrons

**ionic bond** attraction between positively and negatively charged ions

**metallic bond** forms when many metal atoms share their pooled valence electrons